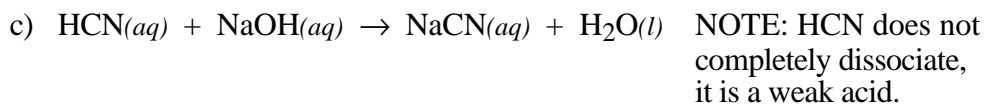
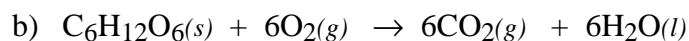
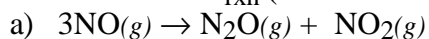
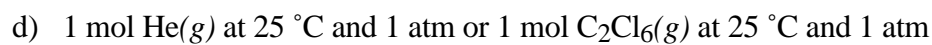
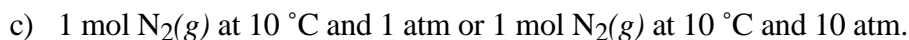
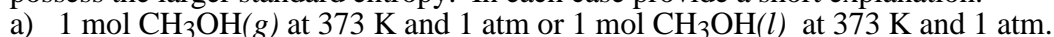


ALL work must be shown to receive full credit. **Due in lecture, at 8:30 a.m. on Friday, January 25, 2002.**

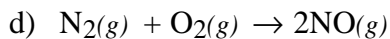
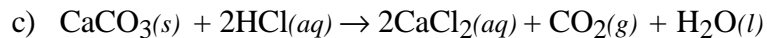
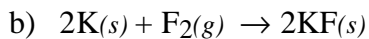
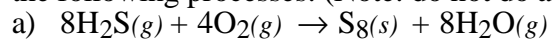
PS1.1. Using the table of thermodynamic values found in Appendix B in your textbook, calculate the  $\Delta H^\circ_{\text{rxn}}$  (standard enthalpy change) for each of the following reactions:



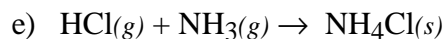
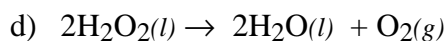
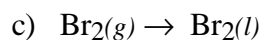
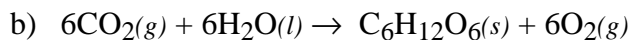
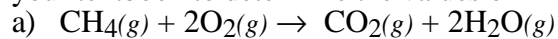
PS1.2. For each of the following pairs, indicate which substance you would expect to possess the larger standard entropy. In each case provide a short explanation.



PS1.3. Predict whether the entropy change in the system is positive or negative for each of the following processes: (Note: do not do any calculations.)



PS1.4. For each reaction below, use the table of thermodynamic values from Appendix B in your textbook to determine the values of  $\Delta H^\circ$  and  $\Delta S^\circ$ .



PS1.5. a) Calculate  $\Delta G^\circ$  for each of the reactions in problem PS1.4.

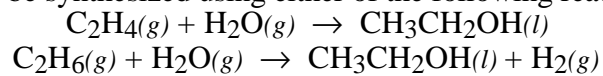
b) Which of the reactions in PS1.4 are spontaneous at 298 K?

c) For each of the reactions listed in b), find the temperature above or below which the reaction becomes nonspontaneous.

d) Which of the reactions in PS1.4 are nonspontaneous at 298 K?

e) For each of the reactions listed in d), find the temperature above or below which the reaction becomes spontaneous.

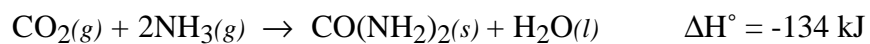
PS1.6. Ethanol could be synthesized using either of the following reactions;



Using thermodynamics arguments which reaction is more feasible under standard conditions? Explain.

PS1.7. When liquid water is introduced into an evacuated vessel at 25 °C, some of the water vaporizes. Predict how the enthalpy, entropy, and free energy change in the system during this process. Explain the reasoning behind your predictions.

PS1.8. Consider the reaction represented below, which is spontaneous at 298 K.



- a) For the reaction, indicate whether the standard entropy change,  $\Delta S^\circ_{\text{rxn}}$  is positive, negative or zero. Justify your answer. (NOTE: do not use a calculation to justify your answer.)
- b) Which factor, the change in enthalpy,  $\Delta H^\circ$ , or the change in entropy,  $\Delta S^\circ$ , provides the principal driving force for the reaction at 298 K? Explain.
- c) For the reaction, how is the value of the standard free energy change,  $\Delta G^\circ$ , affected by an increase in temperature? Explain.

PS1.9. The enthalpy of combustion,  $\Delta H^\circ_{\text{comb}}$ , for oxalic acid,  $\text{C}_2\text{H}_2\text{O}_4(s)$ , is  $-246.05 \text{ kJ}\cdot\text{mol}^{-1}$  and

Substance	$\Delta H^\circ_f \left( \frac{\text{kJ}}{\text{mol}} \right)$	$S^\circ \left( \frac{\text{J}}{\text{mol}\cdot\text{K}} \right)$
$\text{C}(s)$	0	5.69
$\text{CO}_2(g)$	-393.5	213.6
$\text{H}_2(g)$	0	130.6
$\text{H}_2\text{O}(l)$	-285.8	69.96
$\text{O}_2(g)$	0	205
$\text{C}_2\text{H}_2\text{O}_4(s)$	?	120.1

- a) Write the balanced chemical equation that describes the combustion of one mole of oxalic acid.
- b) Write the balanced chemical equation that describes the standard formation of oxalic acid.
- c) Using the information given above and the equations in a) and b), calculate  $\Delta H^\circ_f$  for oxalic acid.
- d) Calculate  $\Delta S^\circ_f$  for oxalic acid and  $\Delta S^\circ_{\text{rxn}}$  for the combustion of one mole of oxalic acid.

e) Calculate  $\Delta G^\circ_f$  for oxalic acid and  $\Delta G^\circ_{\text{rxn}}$  for the combustion of one mole of oxalic acid.

f) Is the formation of oxalic acid from its elements spontaneous? Is the combustion of oxalic acid at 25 °C spontaneous?