A. Acidic Solution	E^o(V)
$F_2(g) + 2H^+ + 2e^- \rightarrow 2HF(aq)$	3.06
$F_2(g) + 2e^- \rightarrow 2F^-$	2.87
$O_3(g) + 2H^+ + 2e^- \rightarrow O_2(g) + H_2O(l)$	2.07
$S_2O_8^{2-} + 2e^- \rightarrow 2SO_4^{2-}$	2.00
$H_2O_2(aq) + 2H^+ + 2e^- \rightarrow 2H_2O(l)$	1.776
$HClO_2(aq) + 2H^+ + 2e^- \rightarrow HClO(aq) + H_2O(l)$	1.645
$2HClO(aq) + 2H^+ + 2e^- \rightarrow Cl_2(g) + 2H_2O(l)$	1.63
$MnO_4^- + 8H^+ + 5e^- \rightarrow Mn^{2+} + 4H_2O(l)$	1.51
$Au^{3+} + 3e^- \rightarrow Au(s)$	1.498
$PbO_2(s) + 4H^+ + 2e^- \rightarrow Pb^{2+} + 2H_2O(l)$	1.455
$Cl_2(g) + 2e^- \rightarrow 2Cl^-$	1.360
$Cr_2O_7^{2-} + 14H^+ + 6e^- \rightarrow 2Cr^{3+} + 7H_2O(l)$	1.33
$2HNO_2(aq) + 4H^+ + 4e^- \rightarrow N_2O(g) + 3H_2O(l)$	1.29
$MnO_2(s) + 4H^+ + 2e^- \rightarrow Mn^{2+} + 2H_2O(l)$	1.23
$O_2(g) + 4H^+ + 4e^- \rightarrow 2H_2O(l)$	1.229
$ClO_3^- + 3H^+ + 2e^- \rightarrow HClO_2(aq) + H_2O(l)$	1.21
$Pt^{2+} + 2e^- \rightarrow Pt(s)$	~1.2
$2IO_3^- + 12H^+ + 10e^- \rightarrow I_2(s) + 6H_2O(l)$	1.195
$ClO_4^- + 2H^+ + 2e^- \rightarrow ClO_3^- + H_2O(l)$	1.19
$Br_2(aq) + 2e^- \rightarrow 2Br^-$	1.087
$Pd^{2+} + 2e^- \rightarrow Pd(s)$	0.987
$NO_3^- + 4H^+ + 3e^- \rightarrow NO(g) + 2H_2O(l)$	0.96
$NO_3^- + 3H^+ + 2e^- \rightarrow HNO_2(aq) + H_2O(l)$	0.942
$2Hg^{2+} + 2e^- \rightarrow Hg_2^{2+}$	0.920
$O_2(g) + 4H^+(pH = 7) + 4e^- \rightarrow 2H_2O(l)$	0.83
$2NO_3^- + 4H^+ + 2e^- \rightarrow N_2O_4(g) + 2H_2O(l)$	0.803
$Ag^+ + e^- \rightarrow Ag(s)$	0.799
$Hg^{2+} + 2e^- \rightarrow Hg(l)$	0.788
$Fe^{3+} + e^- \rightarrow Fe^{2+}$	0.771
$O_2(g) + 2H^+ + 2e^- \rightarrow H_2O_2(aq)$	0.682
$MnO_4^- + e^- \rightarrow MnO_4^{2-}$	0.564
$I_2(s) + 2e^- \rightarrow 2I^-$	0.536
$H_2SO_3(aq) + 4H^+ + 4e^- \rightarrow S(s) + 3H_2O(l)$	0.450
$SO_4^{2-} + 8H^+ + 6e^- \rightarrow S(s) + 4H_2O(l)$	0.357
$Cu^{2+} + 2e^- \rightarrow Cu(s)$	0.337
$AgCl(s) + e^- \rightarrow Ag(s) + Cl^-$	0.222
$SO_4^{2-} + 4H^+ + 2e^- \rightarrow H_2SO_3(aq) + H_2O(l)$	0.172
$Cu^{2+} + e^- \rightarrow Cu^+$	0.153
$Sb_2O_3(s) + 6H^+ + 6e^- \rightarrow 2Sb(s) + 3H_2O(l)$	0.152
$Sn^{4+} + 2e^- \rightarrow Sn^{2+}$	0.15
$S(s) + 2H^+ + 2e^- \rightarrow H_2S(aq)$	0.142
$2H^+ + 2e^- \rightarrow H_2(g)$	0.000
$Pb^{2+} + 2e^- \rightarrow Pb(s)$	−0.126
$Sn^{2+} + 2e^- \rightarrow Sn(s)$	−0.136
$Ni^{2+} + 2e^- \rightarrow Ni(s)$	−0.250
$PbCl_2(s) + 2e^- \rightarrow Pb(s) + 2Cl^-$	−0.268
$PbSO_4(s) + 2e^- \rightarrow Pb(s) + SO_4^{2-}$	−0.359
$Cd^{2+} + 2e^- \rightarrow Cd(s)$	−0.403
$Cr^{3+} + e^- \rightarrow Cr^{2+}$	−0.408
$Fe^{2+} + 2e^- \rightarrow Fe(s)$	−0.440
$Cr^{3+} + 3e^- \rightarrow Cr(s)$	−0.744
$Zn^{2+} + 2e^- \rightarrow Zn(s)$	−0.763
$Mn^{2+} + 2e^- \rightarrow Mn(s)$	−1.185
$Al^{3+} + 3e^- \rightarrow Al(s)$	−1.662
$H_2(g) + 2e^- \rightarrow 2H^-$	−2.25
$Mg^{2+} + 2e^- \rightarrow Mg(s)$	−2.363
$Na^+ + e^- \rightarrow Na(s)$	−2.714
$Ca^{2+} + 2e^- \rightarrow Ca(s)$	−2.866
$Sr^{2+} + 2e^- \rightarrow Sr(s)$	−2.888
$Ba^{2+} + 2e^- \rightarrow Ba(s)$	−2.906
$K^+ + e^- \rightarrow K(s)$	−2.925
$Li^+ + e^- \rightarrow Li(s)$	−3.045

B. Alkaline Solution	E^o(V)
$O_3(g) + H_2O(l) + 2e^- \rightarrow O_2(g) + 2OH^-$	1.24
$ClO^- + H_2O(l) + 2e^- \rightarrow Cl^- + 2OH^-$	0.89
$MnO_4^- + 2H_2O(l) + 3e^- \rightarrow MnO_2(s) + 4OH^-$	0.588
$NiO_2(s) + 2H_2O(l) + 2e^- \rightarrow Ni(OH)_2(s) + 2OH^-$	0.490
$O_2(g) + 2H_2O(l) + 4e^- \rightarrow 4OH^-$	0.401
$ClO_4^- + H_2O(l) + 2e^- \rightarrow ClO_3^- + 2OH^-$	0.36
$ClO_3^- + H_2O(l) + 2e^- \rightarrow ClO_2^- + 2OH^-$	0.33
$PbO_2(s) + H_2O(l) + 2e^- \rightarrow PbO(s) + 2OH^-$	0.247
$HgO(s) + H_2O(l) + 2e^- \rightarrow Hg(l) + 2OH^-$	0.098
$NO_3^- + H_2O(l) + 2e^- \rightarrow NO_2^- + 2OH^-$	0.01
$MnO_2(s) + 2H_2O(l) + 2e^- \rightarrow Mn(OH)_2(s) + 2OH^-$	−0.05
$CrO_4^{2-} + 4H_2O(l) + 3e^- \rightarrow Cr(OH)_3(s) + 5OH^-$	−0.13
$Cu_2O(s) + H_2O(l) + 2e^- \rightarrow 2Cu(s) + 2OH^-$	−0.358
$2S(s) + 2e^- \rightarrow S_2^{2-}$	−0.447
$Fe(OH)_3(s) + e^- \rightarrow Fe(OH)_2(s) + OH^-$	−0.56
$BrO_3^- + 3H_2O(l) + 6e^- \rightarrow Br^- + 6OH^-$	0.61
$Ni(OH)_2(s) + 2e^- \rightarrow Ni(s) + 2OH^-$	−0.72
$Cd(OH)_2(s) + 2e^- \rightarrow Cd(s) + 2OH^-$	−0.809
$2H_2O(l) + 2e^- \rightarrow H_2(g) + 2OH^-(pH = 7)$	−0.43
$2H_2O(l) + 2e^- \rightarrow H_2(g) + 2OH^-$	−0.828
$Fe(OH)_2(s) + 2e^- \rightarrow Fe(s) + 2OH^-$	−0.877
$Zn(OH)_2(s) + 2e^- \rightarrow Zn(s) + 2OH^-$	−1.245
$Al(OH)_3(s) + 3e^- \rightarrow Al(s) + 3OH^-$	−2.30
$Mg(OH)_2(s) + 2e^- \rightarrow Mg(s) + 2OH^-$	−2.690