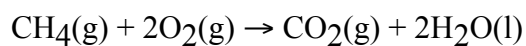
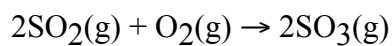
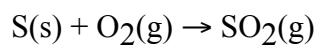


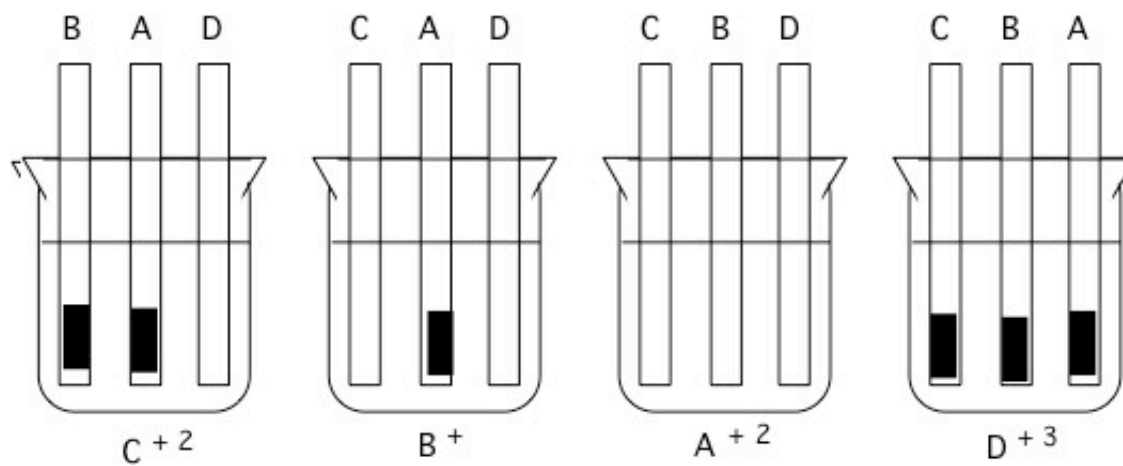
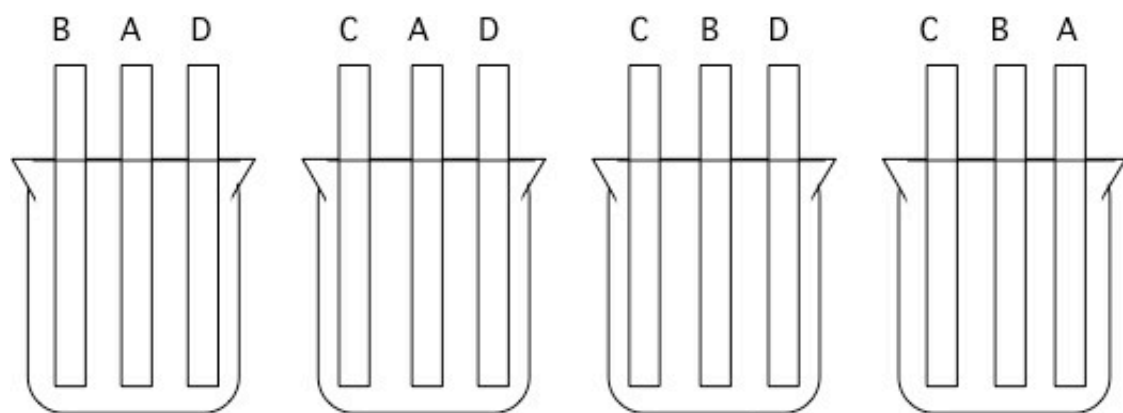
# AP<sup>®</sup> REVIEW ELECTROCHEMISTRY

## Recognizing Oxidation-Reduction Reactions



Identify two characteristics common to these equations.

Consider the following two Figures



Answer the following questions.

What is the most reactive metal ion? The least reactive metal ion?

**What is the least reactive metal? The most reactive metal?**

**Place the other two metal ions and metals.**

**What is interesting about the most reactive metal ion and the least reactive metal?**

**What is a half-reaction?**

**When the half-reactions are arranged in this way we call it an Activity Series for the four metal ions and metals.**

**What pattern do you see in the reactivity of a particular metal and the metal ions.**

**Consider the position of the reactive and nonreactive metal ions to make a rule based on the position of the metal and the metal ions in the activity series that allows one to predict what metal ions will react with what metals.**