

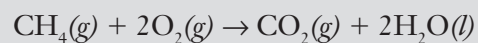
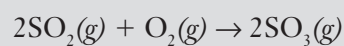
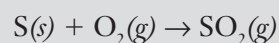
INTRODUCTION TO OXIDATION-REDUCTION

NAME _____

SECTION _____

1. a. Identify each of the chemical equations below as to the reaction type (combustion, formation, or decomposition).

Chemical Equation	Reaction Type
$S(s) + O_2(g) \rightarrow SO_2(g)$	
$2SO_2(g) + O_2(g) \rightarrow 2SO_3(g)$	
$CH_4(g) + 2O_2(g) \rightarrow CO_2(g) + 2H_2O(l)$	
$2HgO(s) \rightarrow 2Hg(l) + O_2(g)$	



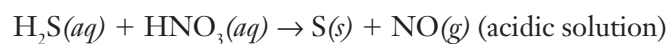
- b. Identify two characteristics common to these equations.

- c. Explain the historic interpretation of oxidation and reduction in chemical reactions.

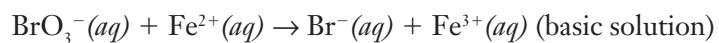
- d. Define the terms oxidation, reduction, oxidizing agent, and reducing agent.

2. Summarize the steps required to balance an oxidation–reduction reaction in aqueous solution.

3. a. Balance the following oxidation–reduction equation:



b. Balance the following oxidation–reduction equation:



c. Balance the following oxidation–reduction equation:

