## NEUTRALIZATION REACTIONS

Name	SECTION

- 1. Write the ionic and net ionic equations for the following reactions, the equilibrium expression for the net ionic equation, and determine the magnitude of the equilibrium constant.
  - a.  $HCl(aq) + NaOH(aq) \rightarrow H_2O(l) + NaCl(aq)$

b.  $HC_2H_3O_2(aq) + KOH(aq) \rightarrow H_2O(l) + KC_2H_3O_2(aq)$ 

c.  $\text{HNO}_3(aq) + \text{NH}_3(aq) \rightarrow \text{NH}_4 \text{NO}_3(aq)$ 

What is the general conclusion which can be made from the magnitude of the equilibrium constant in all three examples above?