

NOMENCLATURE PART II

NAME _____

SECTION _____

1. Below is a list of formulas of ionic compounds. Organize the compounds into sets.

NaCl	$\text{Fe}(\text{NO}_3)_2$	NaI	RbI
KBr	$\text{Fe}(\text{NO}_3)_3$	Li_3PO_4	BaF_2
MgCl_2	KNO_3	Na_2CO_3	CuSO_4
AlCl_3	CaBr_2	CsF	NH_4Cl
AlBr_3	NH_4NO_3	AgCl	PbI_2

2. Complete the following tables:

Name of the Compound	Formula of the Compound	Ionic or Covalent Compound
	$(\text{NH}_4)_2\text{SO}_4$	
Calcium chromate		
	C_6H_{14}	
	PCl_5	
Nitric acid		
	CoCl_3	

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Name of the Compound	Formula of the Compound	Ionic or Covalent Compound
	NH_3	
Potassium phosphate		
	$\text{AgC}_2\text{H}_3\text{O}_2$	
Dichlorine oxide		
	H_2SO_4 (aq)	
Octane		

Name of the Compound	Formula of the Compound	Ionic or Covalent Compound
Sodium sulfide		
Phosphorus pentachloride		
	$\text{Ni}_3(\text{PO}_4)_2$	
Potassium peroxide		
	HNO_3 (aq)	
	C_7H_{16}	