

During Class Invention

Name(s) with Lab section in Group

Nomenclature Part II

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1. Below is a list of formulas of ionic compounds. Organize the compounds into sets.

NaCl	Fe(NO <sub>3</sub> ) <sub>2</sub>	NaI	RbI
KBr	Fe(NO <sub>3</sub> ) <sub>3</sub>	Li <sub>3</sub> PO <sub>4</sub>	BaF <sub>2</sub>
MgCl <sub>2</sub>	KNO <sub>3</sub>	Na <sub>2</sub> CO <sub>3</sub>	CuSO <sub>4</sub>
AlCl <sub>3</sub>	CaBr <sub>2</sub>	CsF	NH <sub>4</sub> Cl
AlBr <sub>3</sub>	NH <sub>4</sub> NO <sub>3</sub>	AgCl	PbI <sub>2</sub>

2. Complete the following table;

Name of the compound	Formula of the compound	Ionic or Covalent Compound
<b>Ammonium sulfate</b>	$(\text{NH}_4)_2\text{SO}_4$	<b>ionic</b>
calcium chromate	<b><math>\text{CaCrO}_4</math></b>	<b>ionic</b>
<b>hexane</b>	$\text{C}_6\text{H}_{14}$	<b>covalent</b>
<b>Phosphorus pentachloride</b>	$\text{PCl}_5$	<b>covalent</b>
nitric acid	<b><math>\text{HNO}_3</math></b>	
<b>Cobalt(III) chloride</b>	$\text{CoCl}_3$	<b>ionic</b>

Name of the compound	Formula of the compound	Ionic or Covalent Compound
<b>ammonia</b>	$\text{NH}_3$	<b>covalent</b>
potassium phosphate	<b><math>\text{K}_3\text{PO}_4</math></b>	<b>ionic</b>
<b>Silver acetate</b>	$\text{AgC}_2\text{H}_3\text{O}_2$	<b>ionic</b>
dichlorine oxide	<b><math>\text{Cl}_2\text{O}</math></b>	<b>covalent</b>
<b>Sulfuric acid</b>	$\text{H}_2\text{SO}_{4(\text{aq})}$	
octane	<b><math>\text{C}_8\text{H}_{18}</math></b>	<b>covalent</b>

Name of the compound	Formula of the compound	Ionic or Covalent Compound
sodium sulfide	<b><math>\text{Na}_2\text{S}</math></b>	<b>ionic</b>
Phosphorus pentachloride	<b><math>\text{PCl}_5</math></b>	<b>covalent</b>
<b>Nickel(II) phosphate</b>	$\text{Ni}_3(\text{PO}_4)_2$	<b>ionic</b>
Potassium peroxide	<b><math>\text{H}_2\text{O}_2</math></b>	<b>covalent</b>
<b>Nitric acid</b>	$\text{HNO}_{3(\text{aq})}$	
<b>octane</b>	$\text{C}_8\text{H}_{18}$	<b>covalent</b>