

## SALTS II

NAME \_\_\_\_\_

SECTION \_\_\_\_\_

1. Calculate the pH of a 0.700 M  $\text{NaC}_2\text{H}_3\text{O}_2$ .  $K_a(\text{HC}_2\text{H}_3\text{O}_2) = 1.8 \times 10^{-5}$ .

2. Calculate the pH of a 0.392 M  $\text{CH}_3\text{NH}_3\text{NO}_3$ .  $K_b(\text{CH}_3\text{NH}_2) = 4.4 \times 10^{-4}$ .