

Standard Reduction Potentials at 25 °C

<u>Half-Reaction</u>	<u>E°</u>
$F_{2(g)} + 2e^- \rightarrow 2F^-(aq)$	+2.87 v
$H_2O_{2(aq)} + 2H^+(aq) + 2e^- \rightarrow 2H_2O(l)$	+1.77 v
$MnO_4^-(aq) + 8H^+(aq) + 5e^- \rightarrow Mn^{2+}(aq) + 4H_2O(l)$	+1.52 v
$Cl_{2(g)} + 2e^- \rightarrow 2Cl^-(aq)$	+1.36 v
$Cr_2O_7^{2-}(aq) + 14H^+(aq) + 6e^- \rightarrow 2Cr^{3+}(aq) + 7H_2O(l)$	+1.33 v
$O_{2(g)} + 4H^+(aq) + 4e^- \rightarrow 2H_2O(l)$	+1.23 v
$Br_{2(l)} + 2e^- \rightarrow 2Br^-(aq)$	+1.07 v
$NO_3^-(aq) + 4H^+(aq) + 3e^- \rightarrow NO(g) + 2H_2O(l)$	+0.96 v
$Ag^+(aq) + 1e^- \rightarrow Ag(s)$	+0.80 v
$I_{2(s)} + 2e^- \rightarrow 2I^-(aq)$	+0.53 v
$Cu^{2+}(aq) + 2e^- \rightarrow Cu(s)$	+0.34 v
$AgCl(s) + 1e^- \rightarrow Ag(s) + Cl^-(aq)$	+0.22 v
$2H^+(aq) + 2e^- \rightarrow H_{2(s)}$	0.00 v
$Pb^{2+}(aq) + 2e^- \rightarrow Pb(s)$	-0.13 v
$Sn^{2+}(aq) + 2e^- \rightarrow Sn(s)$	-0.136 v
$Ni^{2+}(aq) + 2e^- \rightarrow Ni(s)$	-0.25 v
$Cr^{3+}(aq) + 3e^- \rightarrow Cr(s)$	-0.74 v
$Zn^{2+}(aq) + 2e^- \rightarrow Zn(s)$	-0.76 v
$2H_2O(l) + 2e^- \rightarrow H_{2(g)} + 2^-OH(aq)$	-0.83 v
$Mn^{2+}(aq) + 2e^- \rightarrow Mn(s)$	-1.18 v
$Al^{3+}(aq) + 3e^- \rightarrow Al(s)$	-1.66 v
$Mg^{2+}(aq) + 2e^- \rightarrow Mg(s)$	-2.37 v
$Na^+(aq) + 1e^- \rightarrow Na(s)$	-2.71 v
$Ca^{2+}(aq) + 2e^- \rightarrow Ca(s)$	-2.87 v
$K^+(aq) + 1e^- \rightarrow K(s)$	-2.93 v
$Li^+(aq) + 1e^- \rightarrow Li(s)$	-3.05 v