

INTERMOLECULAR ATTRACTIVE FORCES

NAME _____

SECTION _____

1. Sketch the orientations of molecules and/or ions involved in the following intermolecular attractive forces. Include at least one specific example where each attractive force is important. For each one, tell what causes the force and describe its strength relative to the others.
 - a. ion-dipole forces

 - b. dipole-dipole forces

 - c. London dispersion forces

 - d. hydrogen-bonding forces

2. Complete the following table:

System	Primary Intermolecular Force	Sketch of Interaction Between Particles
$\text{CH}_2\text{Cl}_2(l)$		
$\text{NH}_3(l)$		
$\text{SO}_2(l)$		
$\text{KBr}(s)$		
$\text{I}_2(s)$		
$\text{NaCl}(aq)$		
$\text{CH}_3\text{CH}_2\text{OH}(aq)$		