

ELECTRON CONFIGURATION PART II

NAME _____

SECTION _____

- Write the complete electron configuration for:
 - F
 - Si
 - Zn
 - Er
- The maximum number of electrons in:
 - the 2nd shell
 - a d subshell
 - a p orbital
 - 3 p subshell
- Draw an orbital diagram for:
 - nitrogen
 - aluminum
 - chromium

4. What does the term “shield” mean when describing the attraction experienced by an electron in an outer shell?

5. Complete the following table:

Element	Total Number of Electrons	Electron Configuration
Hydrogen		
Lithium		
Beryllium		
Boron		
Carbon		
Nitrogen		
Oxygen		
Fluorine		
Sulfur		
Potassium		
Bromine		